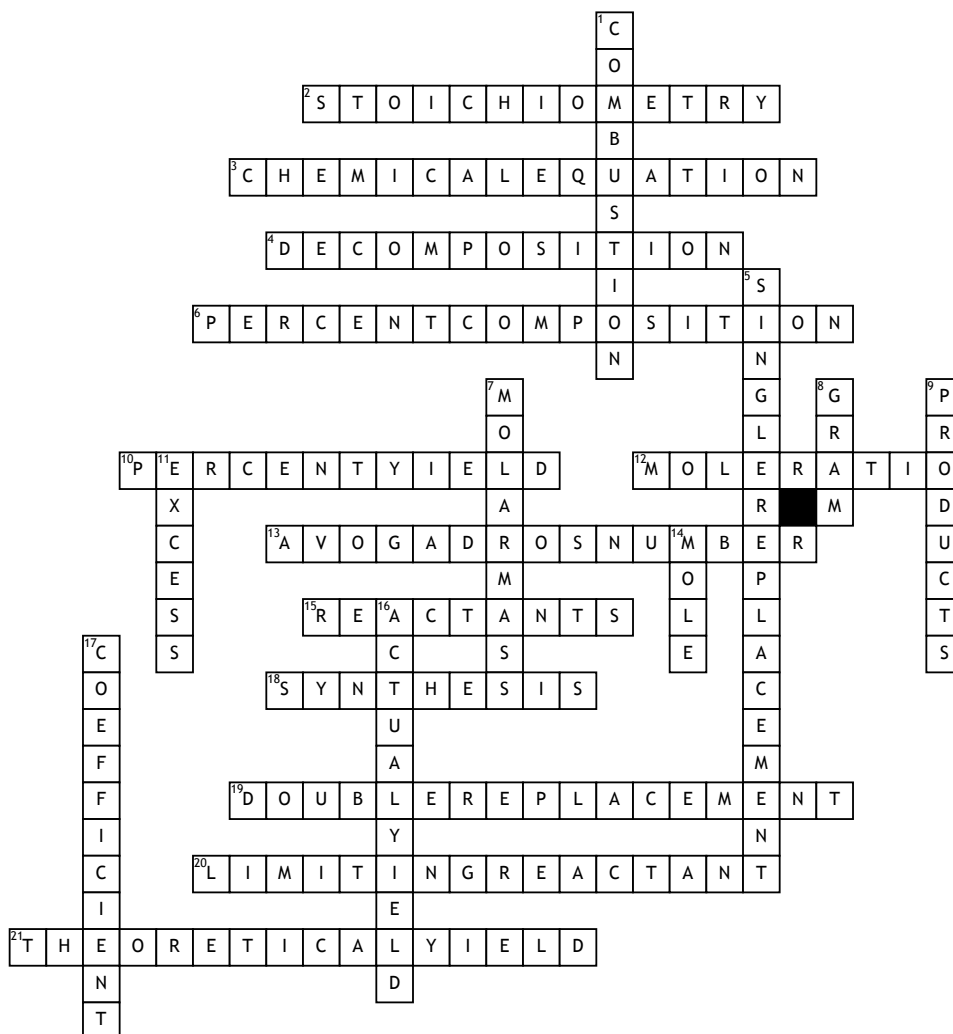


Name: _____

Date: _____

Stoichiometry/Chemical Reactions



Across

2. The area of chemistry involving the calculations of quantities of substances involved in chemical reactions
 3. A representation of a chemical reaction using symbols and numbers to show the relationships between the reactants and products
 4. $AB \rightarrow A + B$
 6. The percent by mass of each element in a compound
 10. The ratio of the actual yield as compared to the theoretical yield expressed as a percentage
 12. A conversion factor derived from the coefficients of the quantities of substances involved in chemical reactions

13. 6.02×10^{23}

15. The chemicals which are put into a chemical reaction and are on the left side of the reaction
 18. $A + B \rightarrow AB$
 19. $AB + CD \rightarrow AD + CB$
 20. The substance that runs out in a chemical reaction, thus controlling the amount of product(s)
 21. The amount of product that could form based on a balanced chemical equation

Down

1. A compound containing carbon and hydrogen is burned in the presence of oxygen
 5. $A + BC \rightarrow AC + B$

7. The mass of one mole of a substance

8. The measurement of the mass of the substances in a chemical reaction
 9. The chemicals present on the right side of a chemical reaction, that are only present after the chemical reaction has begun
 11. The reactant that is leftover after a reaction comes to completion
 14. The amount of a substance
 16. The amount of product that is actually formed when a reaction is carried out in the laboratory
 17. The number in front of a balanced formula showing how much of that reactant or product is present

Word Bank

percent yield	single replacement	decomposition	double replacement	molar mass
theoretical yield	stoichiometry	limiting reactant	reactants	coefficient
actual yield	avogadro's number	chemical equation	gram	excess
combustion	percent composition	synthesis	products	molar ratio
mole				