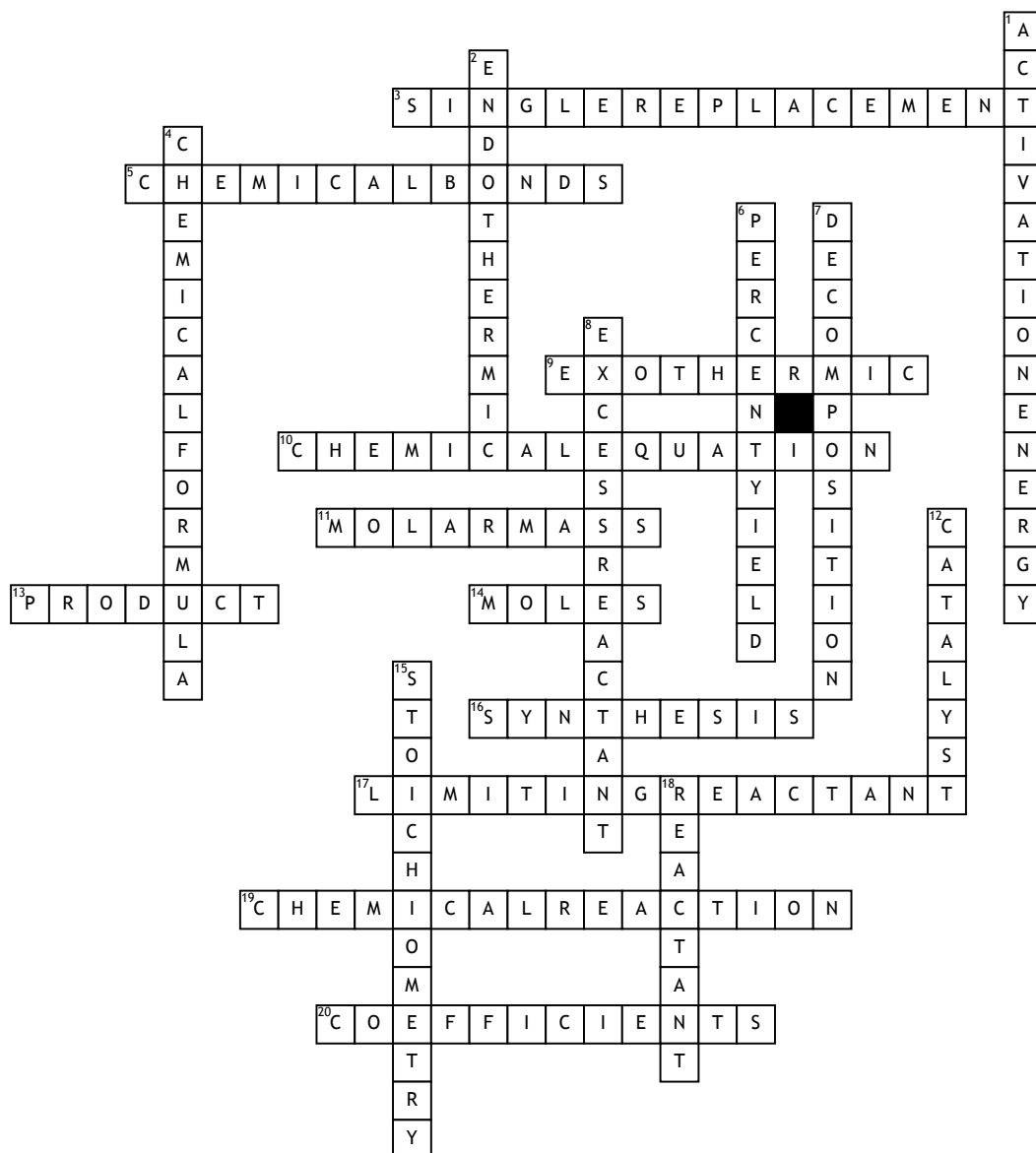


Name: _____

Date: _____

Stoichiometry and Chemical Reactions



Across

3. One element replaces another element in a compound.
 5. What holds individual atoms of compounds together.
 9. A chemical reaction where heat is released to its surroundings.
 10. A representation of a chemical reaction that uses symbols to represent the relationship between the reactants and the products.
 11. The mass of one mole of a substance.
 13. A chemical substance formed as a result of a chemical reaction.
 14. A unit of amount, the amount of a substance containing 6.02×10^{23} entities.

16. Two or more substances combine to form a new compound.

17. A reactant in a chemical reaction that determines the amount of product that is formed.

19. A process in which one or more substances are changed into others.
 20. Used in a chemical equation to determine mole ratios.

Down

1. The smallest amount of energy needed to start a chemical reaction.
 2. A chemical reaction that needs heat.
 4. A combination of chemical symbols and numbers to represent a substance.
 6. The ratio of the actual yield to the theoretical yield shown as a percent.

7. A compound decomposes into two or more simpler substances.

8. The reactant in a chemical equation with a greater amount than necessary to react completely with the limiting reactant.

12. A substance that changes the rate of a chemical reaction without being used up or changed very much.

15. The quantitative relationship between two or more substances especially in processes involving chemical or physical change.

18. A chemical substance that is present at the beginning of a chemical reaction.