

Acid-Base Matching

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| 1. Why is ammonia not considered a base under Arrhenius definition | A. Neutral Salt |
| 2. What is acidic hydrogen | B. It is neutral |
| 3. Why are strong acids and bases good electrolytes | C. To neutralize stomach acids |
| 4. What kind of salt is produced when a strong acid and base react | D. Blue |
| 5. What type of salt is made when a weak base or acid reacts with a strong base or acid | E. 1-7 |
| 6. Why do people use antacid | F. Dissociates to hydroxide ions |
| 7. Why does a small change in pH mean a big change in acidity | G. Because pH is in logs |
| 8. What is the Arrhenius definition of a base | H. Bases |
| 9. What is an Arrhenius acid | I. dissociates to Hydrogen ions |
| 10. Water has a pH of 7 what does this mean | J. Acidic/ Basic salt |
| 11. The pH scale increases from 1-14 , which represents acids | K. Acids |
| 12. A very strong base turns litmus paper | L. Direct reaction with hydrogen |
| 13. What has a low K_b value | M. Hydrogen that produce H^+ ions |
| 14. What has a low K_a value | N. They completely dissociate |