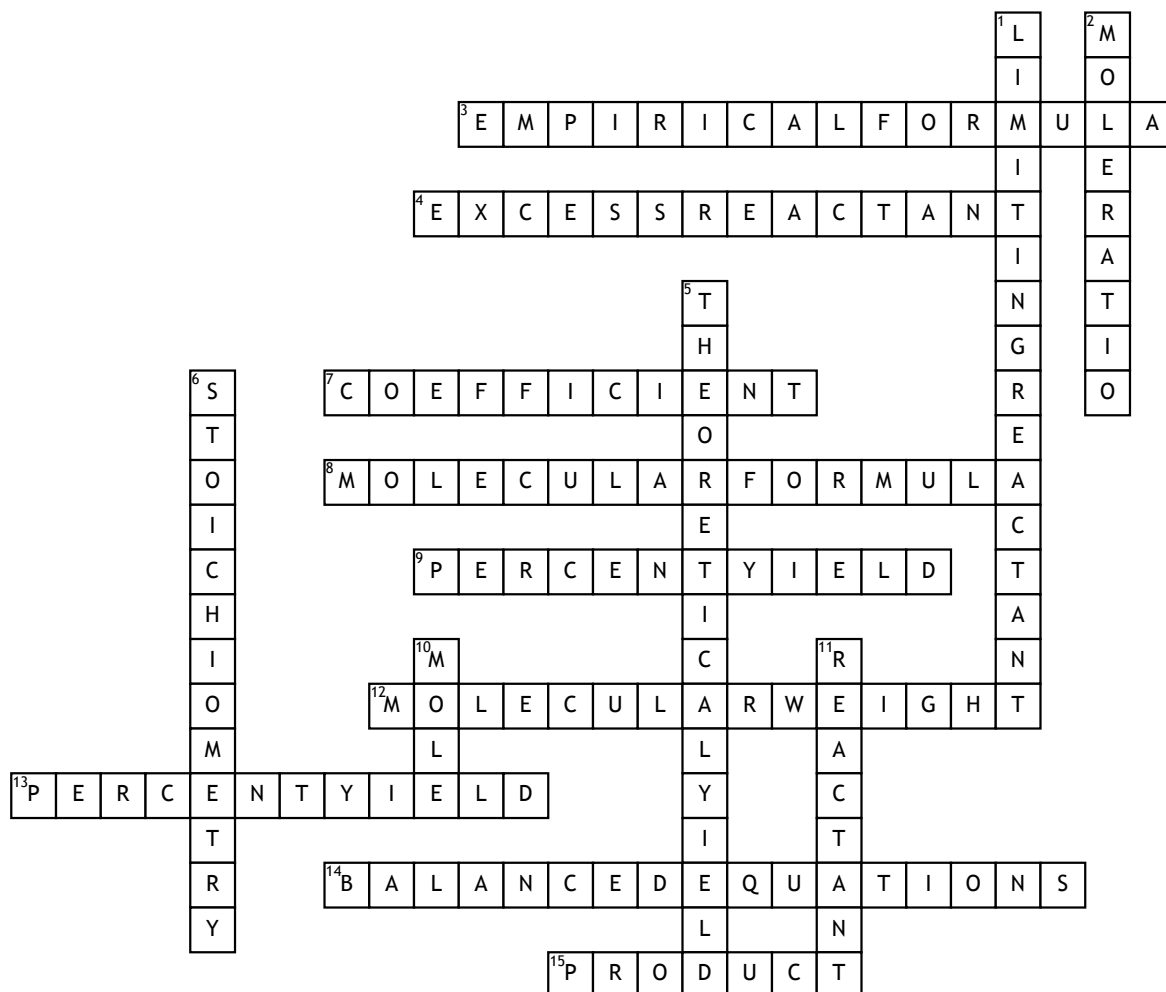


Name: _____

Date: _____

stoichiometry



Across

3. formula using the simplest whole-number ratio
4. reactant that is not used up when a reaction is run to completion
7. a number on the same line of text. In chemistry is the number to the left of a substance in a chemical equation.
8. shows which elements are in the compound and the actual number of each
9. actual yield of a product as a percentage

12. the number of grams in one mole of substance. It is calculated by adding up the atomic masses of all elements in the compound. The units are always g/mol.
13. compares the actual and theoretical yield $(\text{actual}/\text{theoretical}) \times 100$; gives you an idea about the efficiency of a reaction.
14. a chemical reaction that has the same number of atoms of each element on both sides of it; even if they are arranged differently on each side.
15. the substances being made in chemical reaction; the substances on the right side of the arrow in a chemical equation.

Down

1. used up when reaction is run to completion
2. ratio of moles of one substance to another
5. maximum amount of a product that can be given off
6. process of using a balanced equation to determine the relative masses reactants and products involved in the reaction
10. the amount of a pure substance that contains the same number of units
11. the ingredients in a chemical reaction; the substances on the left side of the arrow in a chemical equation.