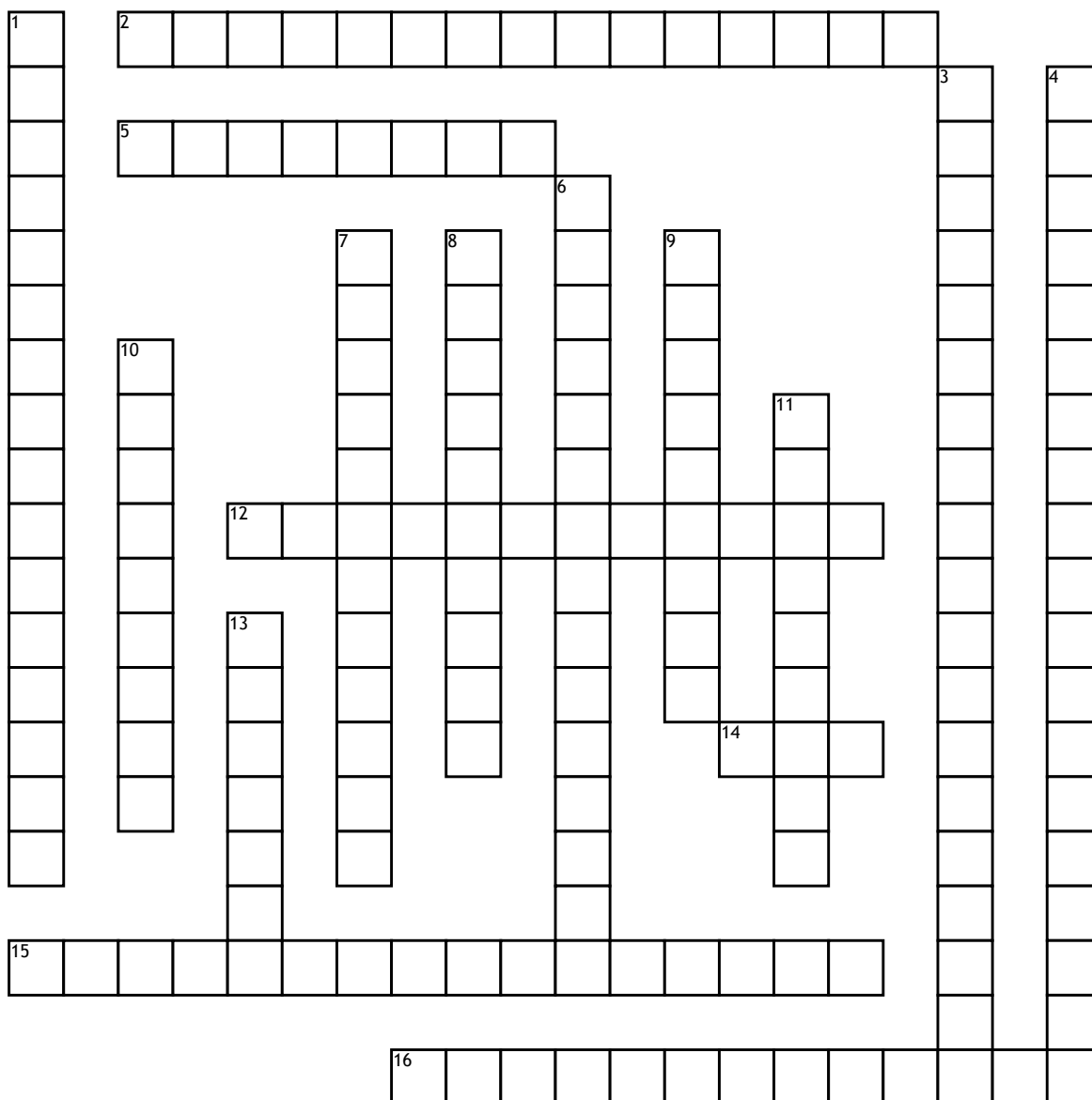


# CHEMICAL REACTIONS VOCABULARY



## Across

**2.** are diagrams that show the bonding between atoms of a molecule and the lone pairs of electrons that may exist in the molecule.

**5.** a substance that takes part in and undergoes change during a reaction.

**12.** a strong force of attraction holding atoms together in a molecule or crystal, resulting from the sharing or transfer of electrons.

**14.** an atom or a molecule in which the total number of electrons is not equal to the total number of protons

**15.** when an electron relocates from an atom to another such chemical entity.

**16.** are a type of chemical bond where two atoms share a pair of electrons with each other.

## Down

**1.** the minimum quantity of energy that the reacting species must possess in order to undergo a specified reaction.

**3.** process describes a process or reaction in which the system absorbs energy from its surroundings; usually, but not always, in the form of heat.

**4.** a chemical reaction that releases energy by light or heat

**6.** a mathematical relationship or rule expressed in symbols.

**7.** a chemical bond that involves the sharing of electron pairs between atoms.

**8.** the state or power of being reactive or the degree to which a thing is reactive.

**9.** a type of chemical bond where a pair of electrons is unequally shared between two atoms.

**10.** a number, figure, symbol, or indicator that is smaller than their normal line of type and is set slightly below or above it.

**11.** the complete transfer of valence electrons between atoms

**13.** a quantity obtained by multiplying quantities together, or from an analogous algebraic operation.