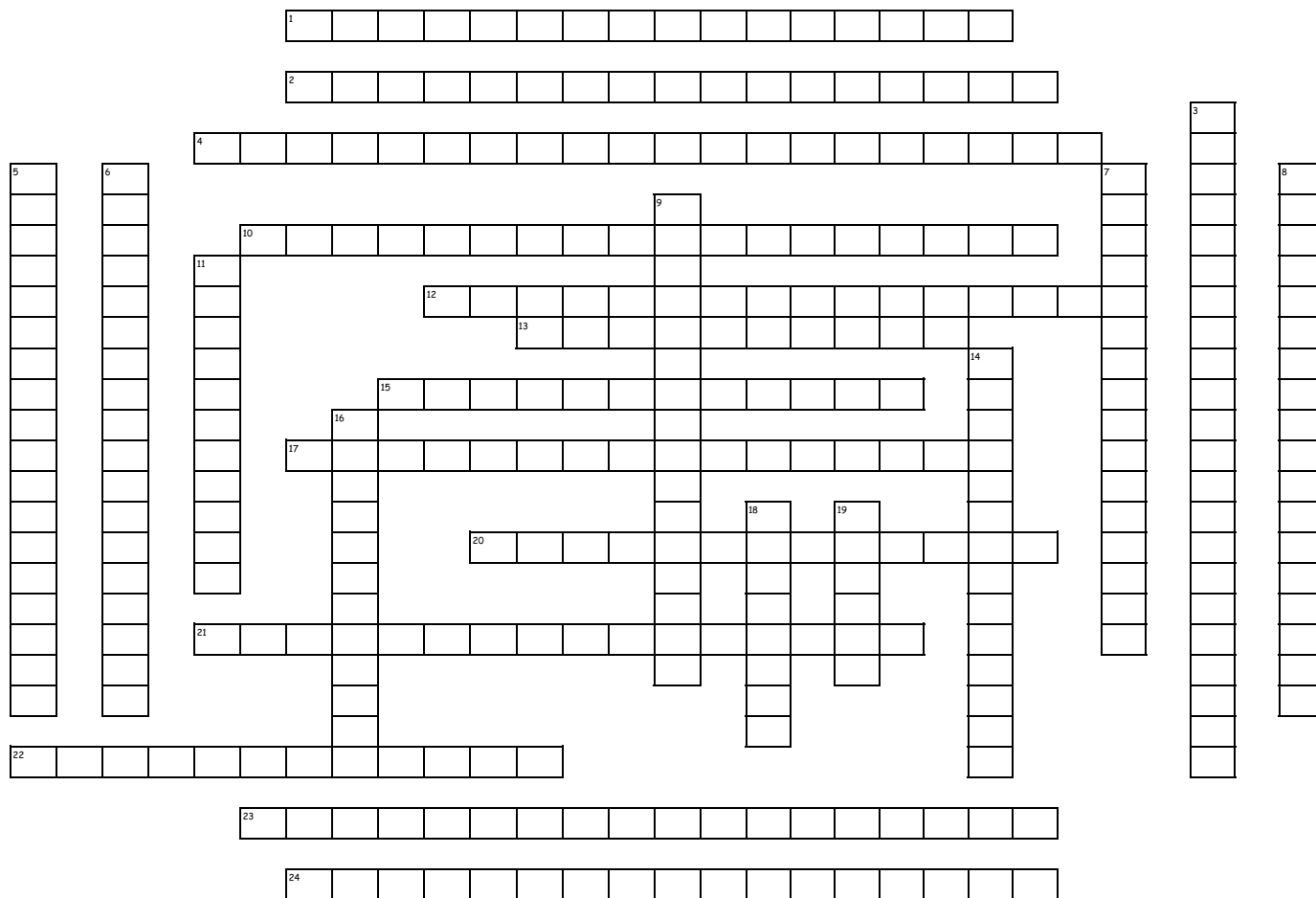


# Covalent Bonding



## Across

1. One measure of the strength of a chemical bond.
2. A formula that shows the arrangement of atoms in the molecule of a compound.
4. Chemical bond where 2 atoms share a pair of electrons.
10. When 2 dipolar molecules interact with each other through space.
12. Attractive or repulsive forces between molecules that don't arise from covalent bonds or ionic bonds
13. Bonds that are partly ionic.
15. A weak bond between 2 molecules.
17. Formula giving the # of atoms in one molecule of a specific compound.
20. Composed of 2 or more atoms covalently bonded.
21. Molecules composed of only 2 atoms.
22. Pair of valence electrons of opposite spin that aren't shared between the atoms in a molecule.
23. 2 pairs of electrons are shared between the atoms rather than just one pair.
24. A compound where the atoms share electrons through covalent bonds.

## Down

3. 2 center, 2 electron covalent bond in which 2 electrons derive from the same atom.
5. A chemical bond between atoms involving 6 bonding electrons instead of 2 in a covalent single bond.
6. A way of describing delocalized electrons within certain molecules where the bonding can't be expressed by one Lewis structure.
7. The weakest intermolecular force.
8. Only one pair of electrons is shared between atoms.
9. A central atom is located at the center with 4 substituents that are located at the corners.
11. A model used in chemistry to predict the geometry of individual molecules from the # of electron pairs surrounding their central atoms.
14. Has a slight positive charge on one side and a slight negative charge on the other.
16. Involves the sharing of electron pairs between atoms.
18. Group of atoms bonded together.
19. A pair of equal and oppositely charged poles separated by a distance.