Name:	Date:

Covalent Bonding

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<u>Across</u>

- 1. One measure of the strength of a chemical bond.
- 2. A formula that shows the arrangement of atoms in the molecule of a compound.
- **4**. Chemical bond where 2 atoms share a pair of electrons.
- 10. When 2 dipolar molecules interact with eachother through space.
- 12. Attractive orrepulsive forces between molecules that dont arise fromcovalent bonds or ionic bonds
- 13. Bonds that are partly ionic.
- 15. A weak bond between 2 molecules.
- 17. Formula giving the # of atoms in one molecule of a specific compound.
- 20. Composed of 2 or more atomscovalently bonded.

- 21. Molecules composed of only 2 atoms.
- 22. Pairof valence electrons of opposite spin that arent shared between the atomsin a molecule.
- 23. 2 pairs of electrons are shared between the atoms rather than just one pair.
- **24**. A compond where the atoms share electrons through covalent bonds.

<u>Down</u>

- 3. 2 center, 2electron covalent bond in which 2 electrons derive from the same atom.
- 5. A chamical bond between atoms involving 6 bonding electrons instead of 2 in a covalent single bond.
- 6. A way of describing delocalized electrons within certain molecules where the bonding cant be expressed by one lewis structure.

- 7. The weakest intermolecular force.
- **8**. Only one pair of electrons is shared between atoms.
- 9. A central atom is located at the center with 4substituentsthat are located at the corners.
- 11. A model used in chemistry to predict the geometry of individual molecules from the # of electron pairs surrounding their central atoms.
- 14. Has a slight positive charge onone side and a slight negative charge on the other.
- **16**. Involves the sharing of electron pairs between atoms.
- 18. Group of atoms bonded together.
- 19. A pair of equal and oppositely charged poles separated by a distance.