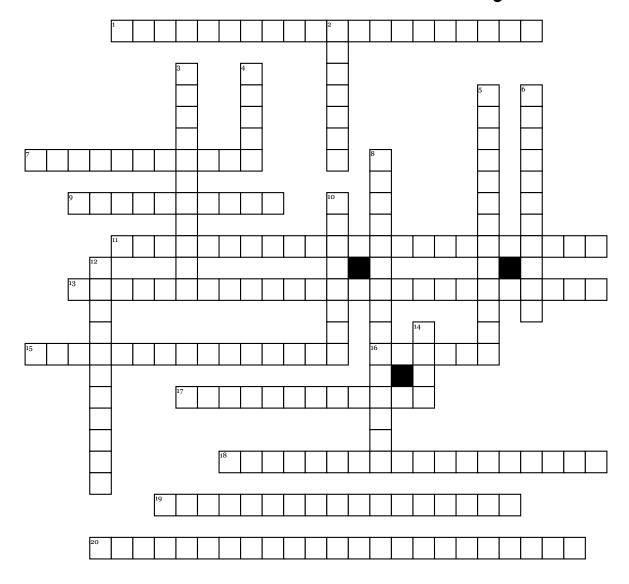
Thermochemistry



Across

- 1. enthalpy change when 1 mole of a compound is completely burnt in Oxygen gas at 298K and 1 bar pressure
- 7. measure of movement of molecules; ONLY movement
- 9. releases heat (feels cold)
- **11.** a quantity that determines transport of matter from one phase to another
- 13. can't create nor destroy energy
- **15.** study of energy transfer in the form of heat during reactions
- **16.** ability to do work (end up somewhere different than where you started)

- 17. number of calories required to raise the temperature of 1 gram of a substance 1 degree Celsius
- **18.** enthalpy change occurring in a system when one mole of matter is changed by a chemical reaction
- **19.** amount of heat necessary to melt (or freeze) **1.00** mole of a substance at its melting point
- **20.** amount of heat necessary to boil (or condense) 1.00 mole of a substance at its boiling point

Down

- 2. an amount of heat exactly
- 4.1840 Joules (Cal)
- 3. absorbs heat (feels hot)
- 4. standard unit of work or energy

- 5. energy due to motion
- **6.** used to measure quantities of heat
- 8. stored/unused energy
- **10.** the internal energy of a system plus the product of the pressure and volume of the system
- **12.** measurement of the quantity of heat exchanged
- **14.** total energy that can be transferred due to molecular movement