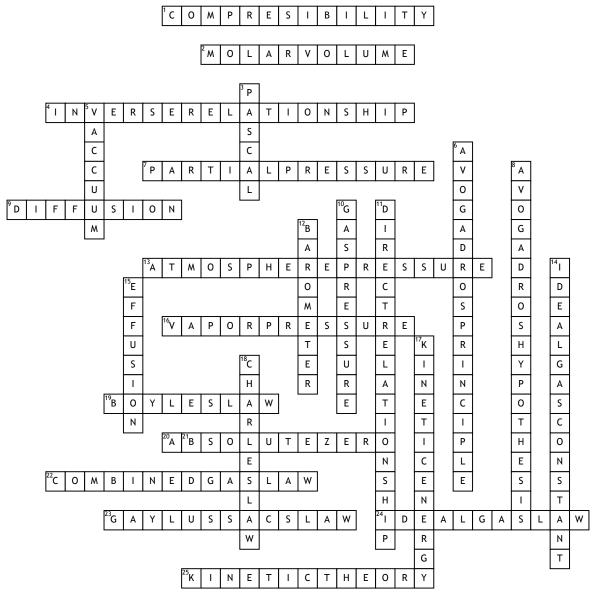
Name:	Date:	

## Gas Laws



## **Across**

- 1. D
- 2. the volume occupied by one mole of ideal gas at STP. Its value is: 22.414 L mol 1.
- 4. Z
- **7.** N
- 9. Z
- **13.** S
- 16. E
- **19.** Boyle's law states that at constant temperature for a fixed mass, the absolute pressure and the volume of a gas are inversely proportional.
- **20.** The lowest temperature that is theoretically possible.

- **22.** When we put Boyle's law, Charles' law, and Gay-Lussac's law together, we come up with the
- **23.** thermal expansion of gasses and the relationship between temperature, volume, and pressure.
- **24.** A physical law describing the relationship of the measurable properties of an ideal gas
- **25.** Z

## **Down**

- **3.** S
- **5.** S
- **6.** states that, "equal volumes of all gases, at the same temperature and pressure, have the same number of molecules"

- **8.** S
- **10.** S
- 11. H
- **12.** A
- **14.** a physical constant which is featured in many fundamental equations in the physical sciences, such as the ideal gas law and the Nernst equation
- **15.** S
- **17.** S
- **18.** Charles's law is an experimental gas law that describes how gases tend to expand when heated.
- 21. Molar