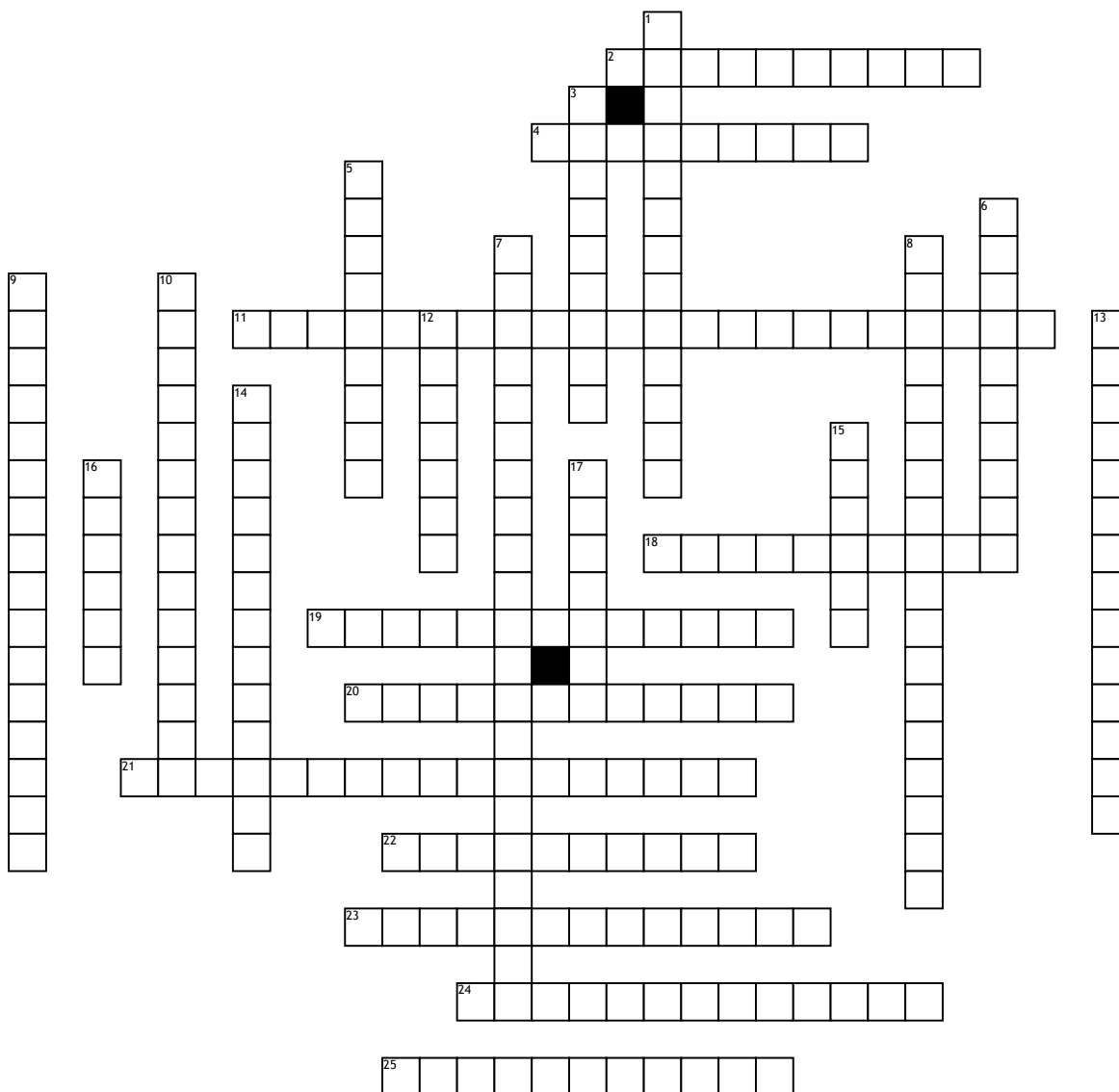


Name: _____

Gas Laws



Across

2. total pressure of a mixture of nonreacting gases is the sum of their individual partial pressures
4. indirect relationship between pressure and volume
11. as particles are heated, they move faster and exert more pressure on the walls of their container
18. Direct relationship between volume and temperature
19. $P_1 \times V_1 = P_2 \times V_2$
20. direct relationship between the number of moles and volume
21. $P_1/T_1 = P_2/T_2$
22. What law happens when you put a balloon into liquid nitrogen?

23. What law is it when you open a sealed container of hot food?
24. What law is it when you heat a can and then put it into water?
25. the ideal law that factors at STP and includes number of moles

Down

1. Direct relationship between pressure and temperature
3. What law is a spray can?
5. What law happens what you twist a straw and then flick it?
6. What law happens what you leave a basketball out side in cold weather for a long time?
7. $P_1V_1/T_1 = P_2V_2/T_2$

8. $PV = nRT$
9. $V_1/n_1 = V_2/n_2$
10. $P_{\text{total}} = P_a + P_b + P_c \dots$
12. Is Boyle's Law a inverse or direct porportion?
13. $V_1/T_1 = V_2/T_2$
14. A law that combines Boyles, Charles, and Lussacs laws
15. Is Charles's law inverse or direct porportion?
16. Is Ideal Gas Law a inverse or direct porportion?
17. What does 1 atm equal?