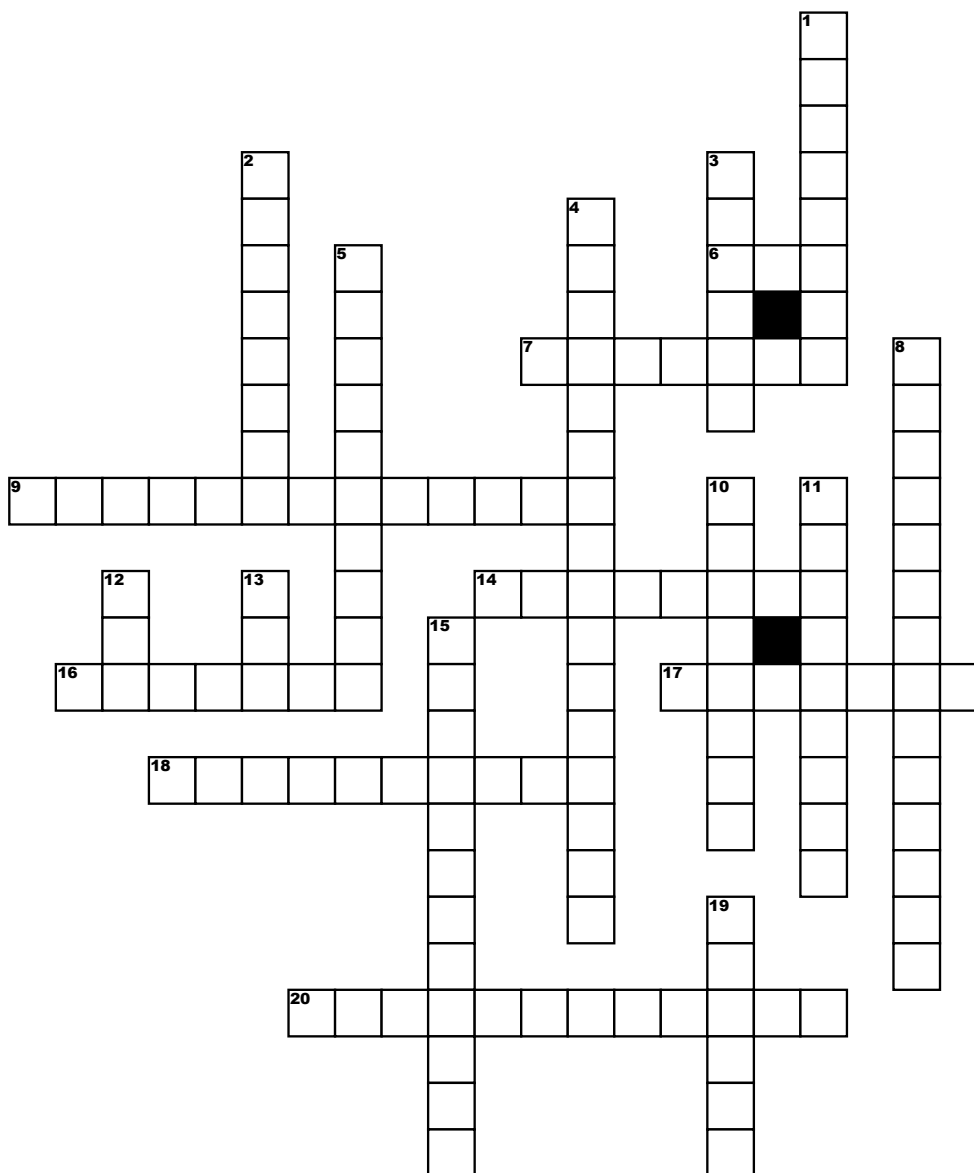


Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

# Structure of the Atom



## **Across**

- 6.** Mass Number - Atomic Number = Number of Neutrons  
**7.** The horizontal rows in the periodic table  
**9.** Arrangement of elements according to atomic number on a table  
**14.** A single atom or several atoms bound together electro magnetically bonded.  
**16.** Center of atom  
**17.** Positive particles around the nucleus  
**18.** The sum of the number of neutrons and protons

- 20.** 1. All matter consists of minuscule particles called atoms.  
**2.** All atoms of a given element are identical to each other. 3. All atoms of a given element are different than those of other elements. 4. Atoms of one element combine with other elements to create compounds. They always combine in equal amounts.

## **Down**

- 1.** Pure substances that can't be broken down or changed  
**2.** 2 or more elements chemically bonded  
**3.** Abbreviation for each element  
**4.** The electrons on the very outer ring of the atom

- 5.** The property of an atom that causes it to have weight

- 8.** Located outside the nucleus with electrons

- 10.** Neutral charged particles around the nucleus

- 11.** Negative particles outside the nucleus

- 12.** Atomic Mass Unit

- 13.** Atomic Number = Protons and Electrons

- 15.** Quantity of protons and electrons in the nucleus of an atom

- 19.** The vertical rows in the periodic table