Across
2. Each __________ on the Periodic Table represents the number of ‘energy levels’ an element has.
4. The elements are organized into this grid.
6. When elements react they form new ________________.
8. The smallest unit of matter with all the properties of that substance.
9. Characteristics that are measurable or observable are called physical __________.
11. An element that can have a variable number of neutrons in its nucleus.
13. The “outer energy shell” of an atom.
15. Each __________ on the periodic table represents the number of valence electrons in an element.
19. A charged particle.
20. A neutron has a ________________ charge.
22. Atoms with full outer energy shells are known to be ____________.
23. The center of an atom where the protons and neutrons are located.
25. Electrons are ‘stolen’ in this type of bond.
26. Reactivity is a ________________ property.
29. An abreviated way to name an element.
30. Subtract the atomic number from the atomic mass to find the number of

Down
1. A “Family” of elements that do not typically react with other elements.
3. The location around the nucleus where electrons orbit.
5. This element is found in all organic matter and has four valence electrons.
7. The ‘electronic connections’ between elements in a molecule.
10. NaCl is the ________________ for salt.
12. Other than hydrogen and helium, the number of electrons needed to fill the valence shell.
13. The “Family” of elements that are very reactive.
14. Protons have a ________________ charge.
17. Electrons have a ________________ charge.
18. ‘Like’ electric charges ________________ each other.
21. A group of elements with similar properties are known as a ________________.
24. This matches the number of protons in an atom.
27. A combination of one or more atoms.
28. Electrons are ‘shared’ in this type of bond.