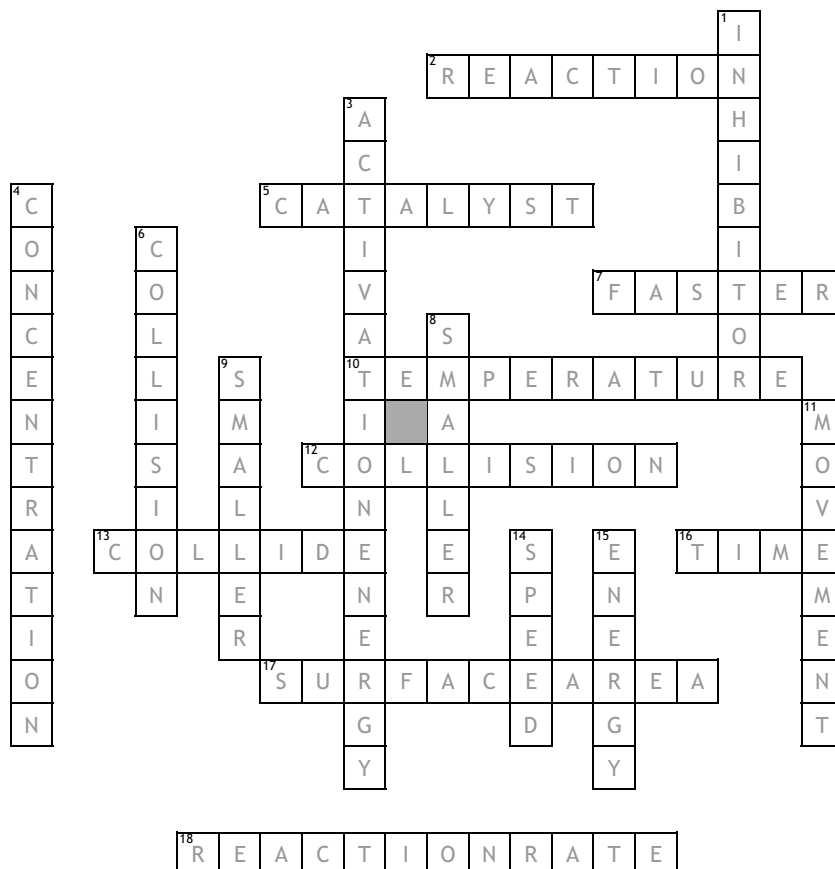


Rates of Reaction



Across

2. Anything that increases the number of successful collisions between reactant particles will speed up a _____
5. The presence of a _____ increases the reaction rate
7. At a higher temperature, particles move _____
10. Increasing the _____ increases the speed of the molecules
12. A _____ reactant will react faster if it is broken into _____ pieces. This increases the _____ and allows more _____ between the reactants
13. Chemicals react together when their particles _____ with each other.
16. The speed of a chemical reaction may be defined as the change in concentration of a substance divided by the _____
17. the smaller a reactant the larger its _____
18. How quickly or slowly a reaction takes place

Down

1. A substance that blocks a catalyst
3. Adding a _____ can also speed up a reaction because it helps to lower the _____
4. At a higher _____, there are more particles in the same amount of space.
6. An increase in surface area means that there will be more _____ between the reactants.
8. A solid reactant will react faster if it is broken into _____ pieces.
9. As the pressure increases, the space in which the gas particles are moving becomes _____
11. More kinetic energy means more _____
14. The rate of reaction is used to describe the _____ of reaction.
15. particles must be moving very fast and have lots of kinetic _____ for collisions to occur.