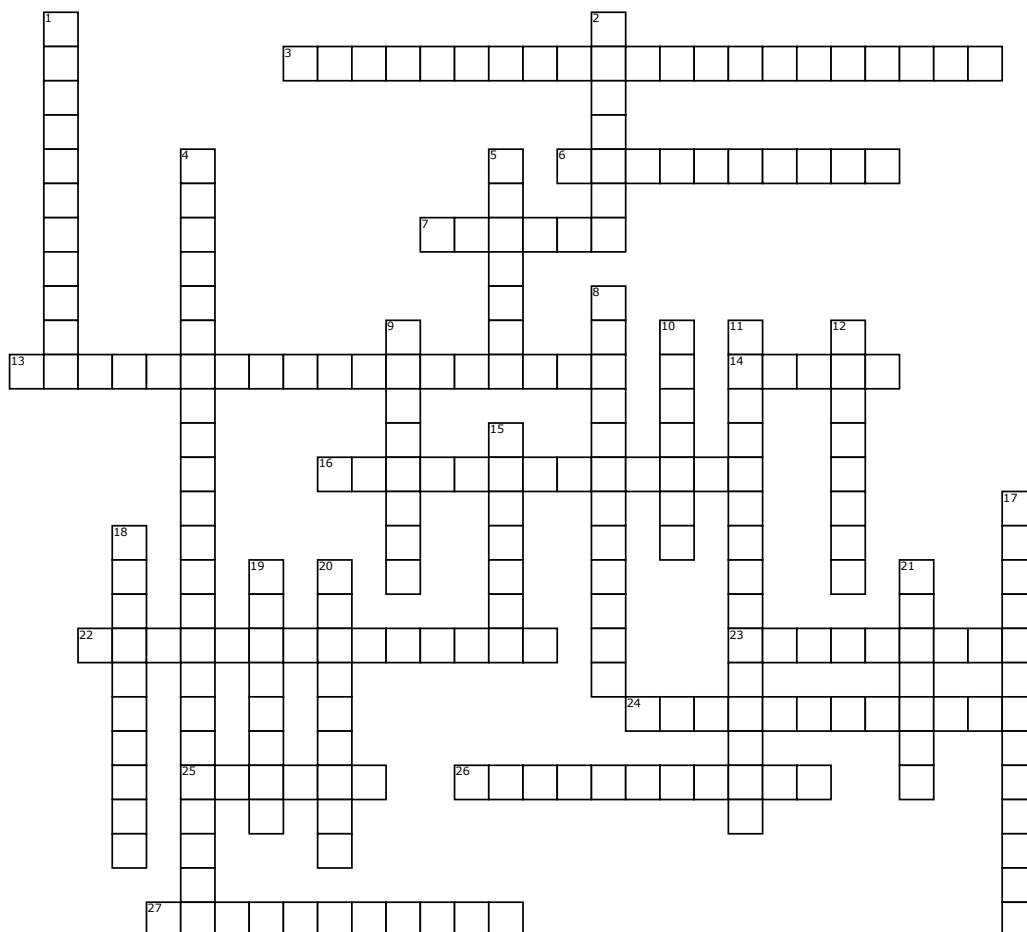


# Regents Chemistry Chapter 12 Solution Vocabulary



## Across

**3.** the boiling point of a solution is higher than the expected boiling point of the pure solvent (colligative property)

**6.** A heterogeneous mixture in which relatively large particles are suspended in a liquid

**7.** having a relatively low concentration of solute in a mixture

**13.** % comp = (part/whole) x 100

**14.** a homogenous mixture/solution containing at least one metal. Ex: brass, steel, bronze

**16.** A sample of matter consisting of more than one pure substance and more than one phase

**22.** a solution in which the concentration of solute is higher than the solubility; more solute is dissolved than should be under a given set of conditions

**23.** Refers to a substance that does not dissolve in a solvent to any significant degree

**24.** the temperature at which a liquid undergoes a phase change from liquid to gas; the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure.

**25.** A substance dissolved in a solvent to make a solution

**26.** A sample of matter consisting of more than one pure substance with properties that do not vary within the sample

**27.** A solution with a concentration lower than its equilibrium solubility; a solution in which more solute can be dissolved

## Down

**1.** An insoluble substance that has formed from a chemical reaction between substances dissolved in a solution

**2.** two or more pure substances PHYSICALLY combined; a combination of two or more pure substances that can be separated by physical means

**4.** the freezing point/melting point of a solution is lower than the freezing point/melting point of the pure solvent (colligative property)

**5.** The most abundant component in a solution

**8.** Having a relatively large amount of substance present in a unit amount of mixture. For example, a 12 M HCL solution is more concentrated than an 0.001 M HCL solution

**9.** a homogeneous mixture

**10.** a heterogeneous mixture composed of tiny particles suspended in another material. The particles are larger than the particles in a solution but smaller than particles in a suspension. Ex: milk, blood

**11.** a measure of concentration; ppm = parts of solute/million parts of solution

**12.** a measure of concentration; M=moles of solute/liters of solution

**15.** capable of being dissolved in a solvent

**17.** A measure of the amount of solute present in a unit amount of mixture. (Ex: ppm = parts per million, molarity = moles solute/liter solution); the process of increasing the amount of substance in a given amount of mixture

**18.** a measure of the concentration of a substance in a saturated solution; a measure of how much a substance can dissolve in a given amount of solvent

**19.** Two liquids are considered "miscible" or mixable if shaking them together results in a single liquid phase with no visible separation

**20.** a solution that has reached equilibrium; a solution which can't dissolve any more solute

**21.** a homogeneous mixture/solution in which a solute is dissolved in water

## Word Bank

Heterogeneous

Concentration

Colloid

Boiling Point

Alloy

Parts Per Million

Mixture

Saturated

Miscible

Insoluble

Solubility

Suspension

Precipitate

Solute

Solvent

Unsaturated

Solution

Soluble

Supersaturated

Percent Composition

Aqueous

Boiling Point Elevation

Freezing Point Depression

Molarity

Concentrated

Homogeneous

Dilute