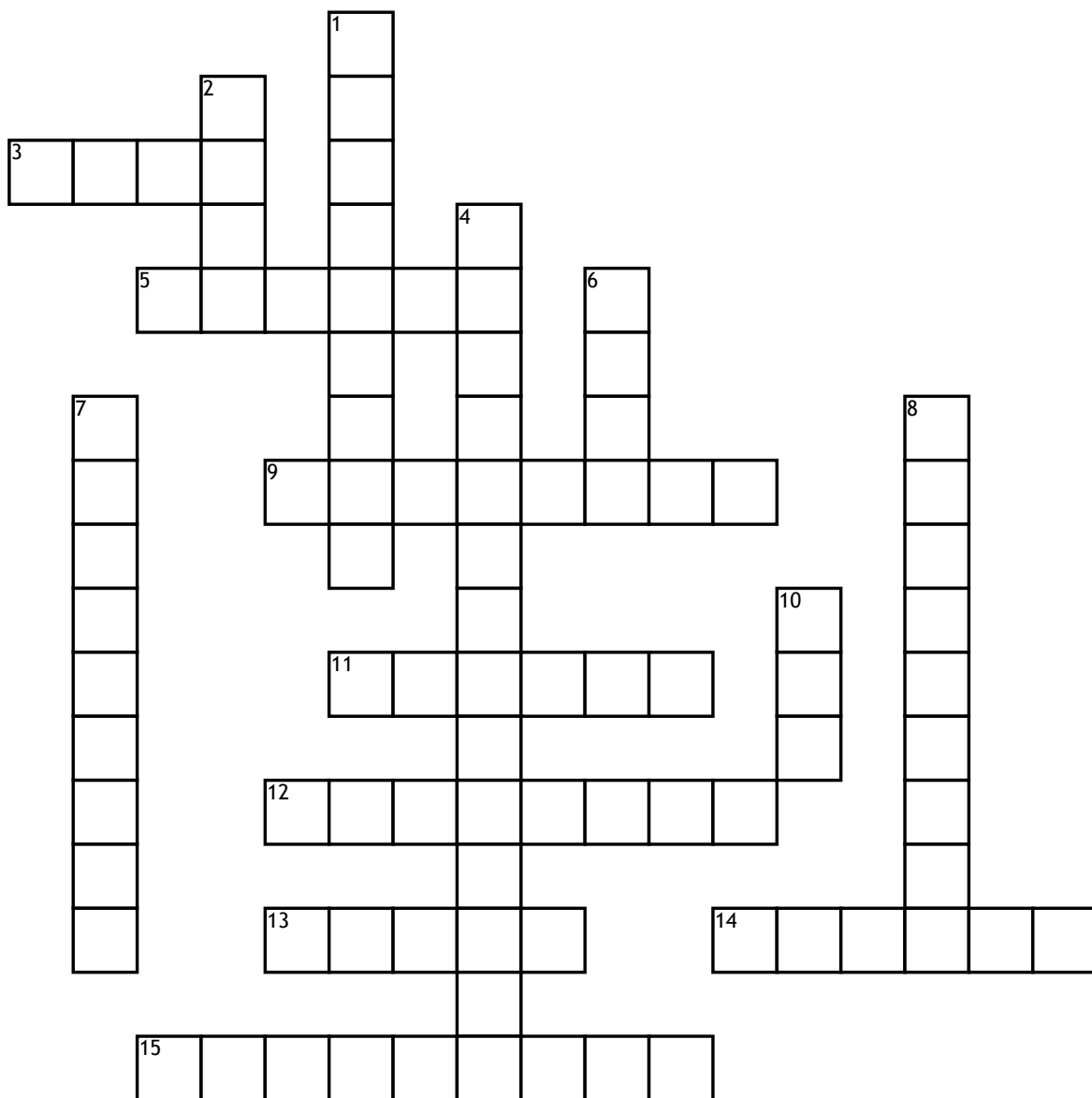


Name: \_\_\_\_\_ Date: \_\_\_\_\_ Period: \_\_\_\_\_

# Acid-Base Chemistry



## Across

3. Bases turn red indicator paper

5. Bronsted-Lowry definition says that acids \_\_\_\_\_  $H^+$

9. Highest pH possible

11. A(n) \_\_\_\_\_ solution has more hydronium than hydroxide

12. Measured in moles per liter

13. A(n) \_\_\_\_\_ solution has more hydroxide than hydronium

14. Bronsted-Lowry definition says that bases \_\_\_\_\_  $H^+$

15.  $H^+$

## Down

1. Method to determine the concentration of a solution

2. Lowest pH possible

4. Reaction between an acid and base resulting in a salt and a water

6. A pH of 2 is \_\_\_\_\_ acidic than a pH of 6

7. The result of the acid donating its proton is the \_\_\_\_\_ base

8.  $OH^-$

10. Acids turn blue indicator paper