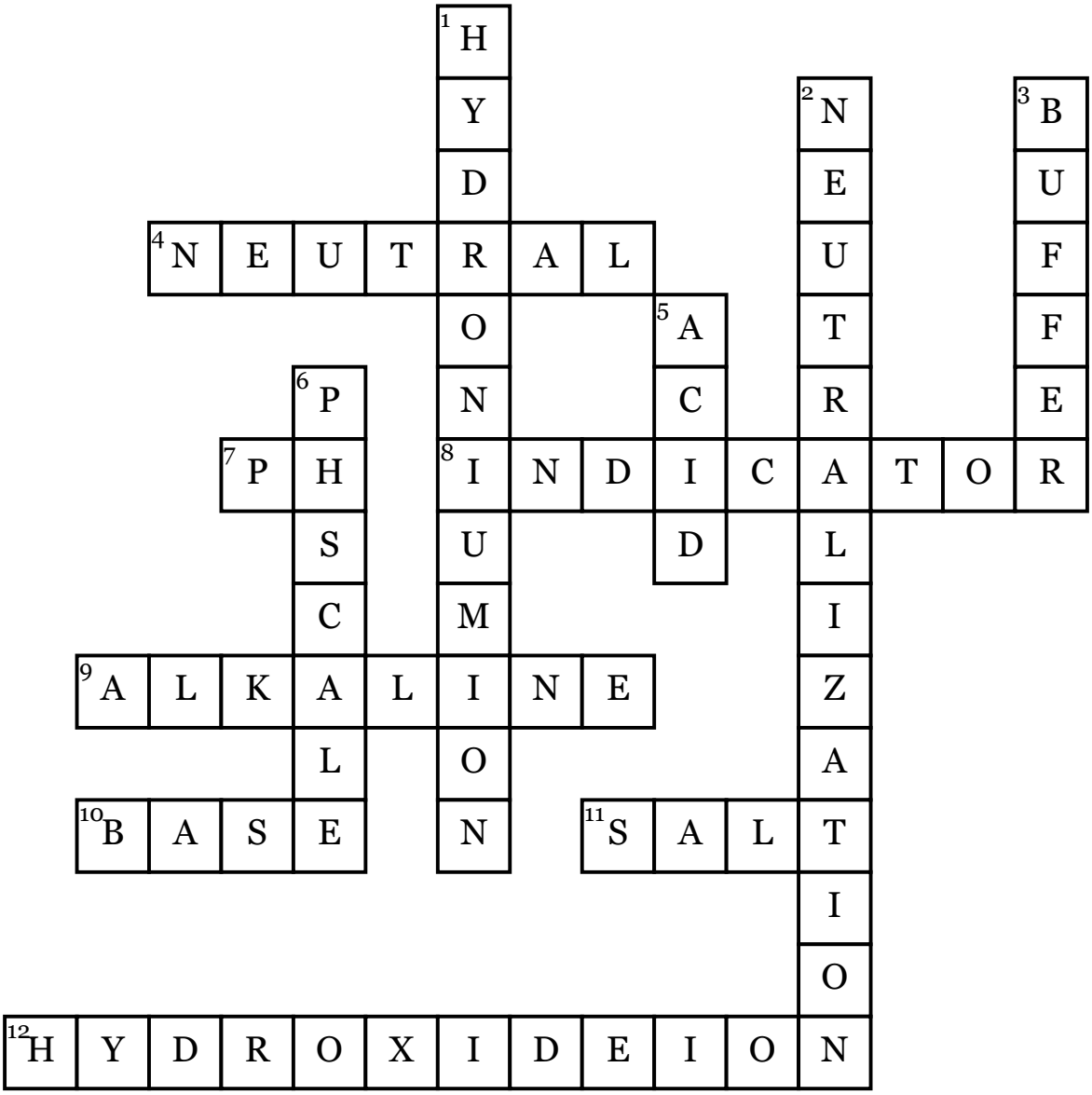
Name: _____ Date: _____

Acid and Bases



Across

4. A solution that is neither acidic nor alkaline, such as pure water.

7. A numeric scale used to specify the acidity or basicity of an aqueous solution.

8. Any substance that gives a visible sign, usually by a colour change, of the presence or absence of a threshold concentration of a chemical species, such as an acid or an alkali in a solution.

9. A chemical compound that neutralizes or effervesces with acids and turns litmus blue; typically.

10. Are substances that, in aqueous solution, are slippery to the touch, taste bitter, change the color of indicators.

11. Any chemical compound formed from the reaction of an acid with a base, with all or part of the hydrogen of the acid replaced by a metal or other cation.

12. The monovalent anion OH– consisting of one atom of hydrogen and one of oxygen.

Down

1. The ion H3O+, consisting of a protonated water molecule and present in all aqueous acids.

2. Is a chemical reaction in which an acid and a base react quantitatively with each other.

3. A solution that resists changes in pH when acid or alkali is added to it. Buffers typically involve a weak acid or alkali together with one of its salts.

5. A molecule or other entity that can donate a proton or accept an electron pair in reactions.

6. A measure of acidity or alkalinity of water soluble substances.