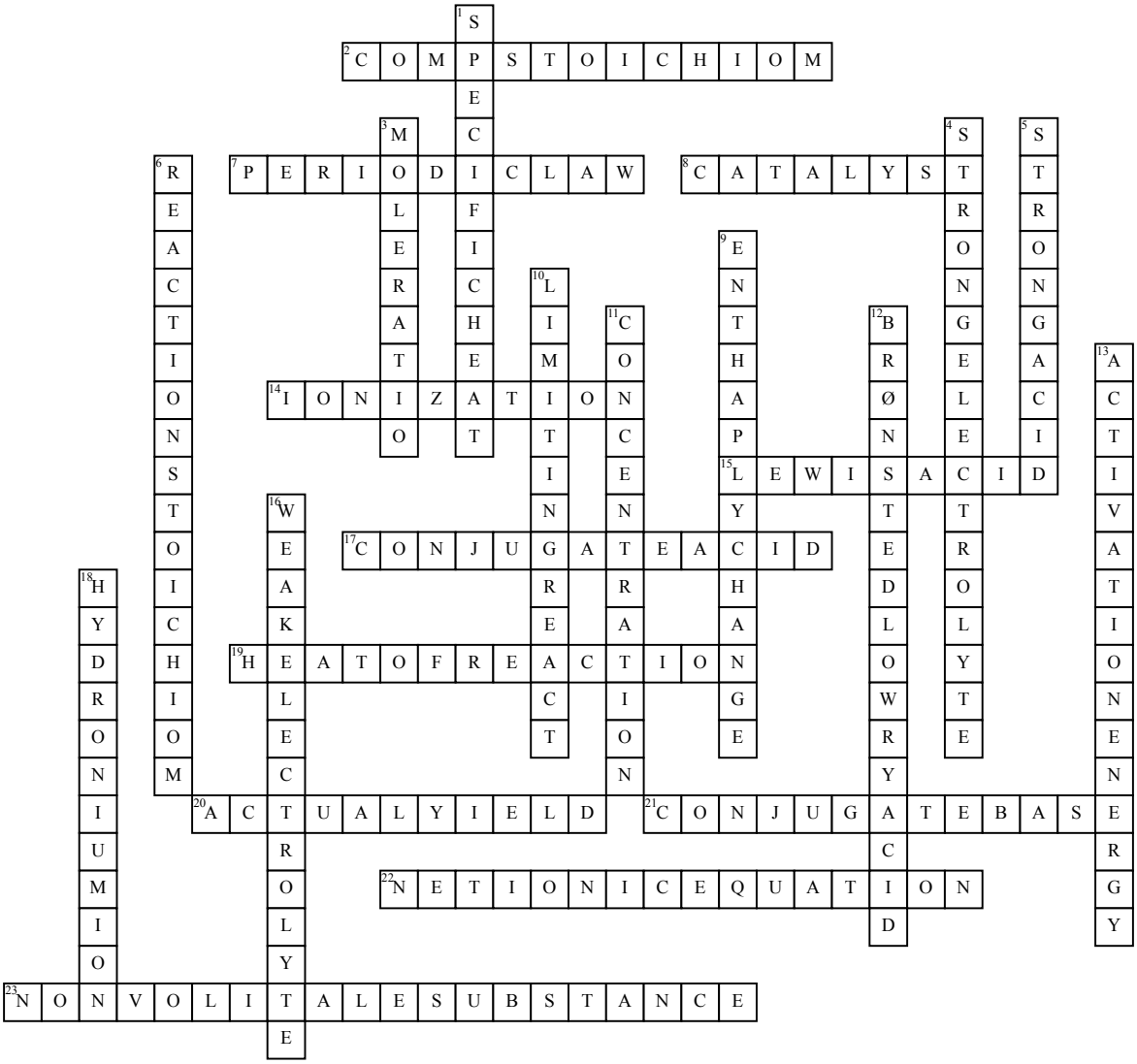
Name: _____ Date: _____ Period: _____

Chemistry Review



Across

2. Chpt.9 page 275- the mass relations of elements

7. Chpt.5 page 125-the physical/ chemical properties of the elements

8. Chpt.17 page 540- is a substance that changes the rate of a chemical reaction without itself being permanently consumed (lowering activation energy)

14. Chpt.14 page 431 - ions are formed from solute molecules by the action of the solvent in this process

15. Chpt 15.page 467 - an ion that accepts an electron pair to form a covalent bond

17. Chpt.15 page 459- species formed when a Brønsted Lowry base gains a proton, that acid base

19. Chpt.17 page 514- quantity of energy released or absorbed as heat during a chemical reaction

20. Chpt.9 page 293- measured product that can be produced from a given amount of reactant

21. Chpt.15 page 469 - species that remains after a Brønsted Lowry acid has given up a proton

22. Chpt.14 page 429-includes only those compounds and ions that undergo a chemical change in a reaction in a aq solution

23. Chpt.14 page 436- one that had little tendency to become a gas under conditions

Down

1. Chpt 17. page 512-amount of energy required to raise the temp of 1g of substance by 1degree Celsius

3. Chpt.9 page 276- relates the amount of moles of any 2 substances

4. Chpt.14 page 432- any compound whose dilute aq solutions conduct electricity well

5. Chpt.15 page 460 - one that ionizes completely in aq solution

6. Chpt.9 page 275 - the relations between Reactants /Products

9. Chpt.17 page 516 -the amount of the energy absorbed or lost by a system as heat during a process at constant pressure

10. Chpt. 9 page 288- that limits the amounts of the other reactants that can combine

11. Chpt .13 page 412 - Measure of amount of Solute in a given amount of solvent/solution

12. Chpt.15 page 464- a molecule that is a proton donor

13. Chpt.17 page 534 - minimum energy required to transform the reactants into an activated complex

16. Chpt.14 page 433-any compound whose dilute aq solutions conduct electricity poorly

18. Chpt 14 page 430- The H30+ ion